Synergy 9

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| Name of Group Members | Solution A | Solution B | Evidence of Chemical RXN? | Mass Change + / - (g) |
| Example and Test | Pb(NO3)2 | KI | Yellow Precipitate Form | + 0.01 |
| Michelle and Ryan | Fe(No3)3 | KSCN | Blood red precipitate form | -0.01 |
| Kevin and Zaiyou | Fe(No3)3 | KSCN | Dark red precipitate form | 0.00 |
| Erica and Margaret | Fe(No3)3 | KSCN | Rusty rouge precipitate form | -0.03 |
| Patrick­ and Patrick2 | CaCl2 | Na2CO3 | White misty precipitate form | +0.02 |
| Ethan and Victor | Fe(No3)3 | KSCN | Bloody precipitate form | +0.05 and then -0.03 |
| Jason Z and Bryan | Pb (NO3)2 | KI | Bright Yellow precipitate formed | ***+9.99*** |
| Amber and Grace | Pb (NO3)2 | KI | Fluorescent Yellow Precipitate Form | 0.00 |
| Nicole and Tania | Pb (NO3)2 | KI | Bright yellow precipitate form | 0.00 |
| Amy Lu and Nicole | Fe(No3)3 | KSCN | Deep Crimson Red Precipitate form | +0.03 |
| C8lyn and Daerile | Pb(NO3)3 | KI | Foamy and yellow | * 0.02 |
| Jason Austin | Fe(NO3)3 | KSCN | It turned dark red | * 0.01 |
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Questions For homework

Record and tabulate all data and observations (qualitative and quantitative) from this lab

1. Why is it important that the flask be sealed for this experiment, even after the flask is returned to an upright position?
2. What observations lead you to believe that a chemical reaction occurred inside the flask?
3. In general, what overall mass change results from a chemical reaction?
4. Suppose that a reaction was carried out in an open flask and the final mass was significantly greater than the initial mass, what could you conclude?
5. If a reaction was carried out in an open flask and the final mass was significantly less than the initial mass, what could you conclude?
6. What is “the law of conservation of mass” for chemical reactions? If necessary use the web or reference texts for help.
7. Draw and label a diagram of the set up that you used today.